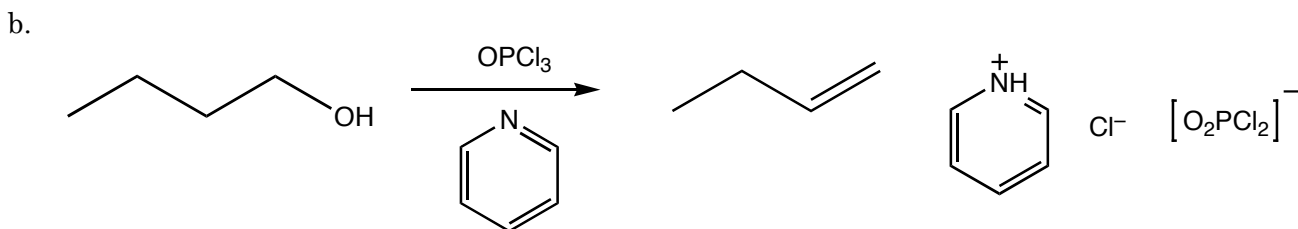
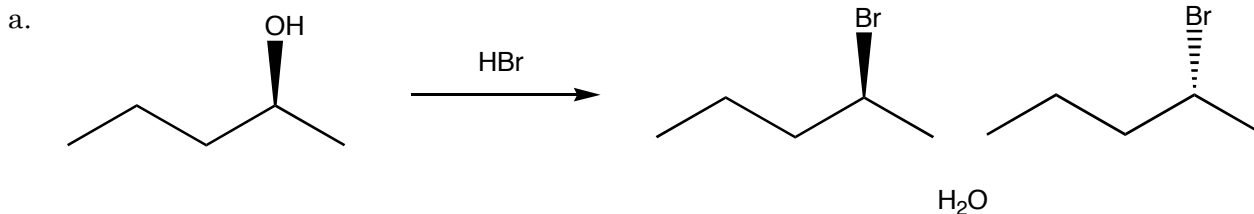
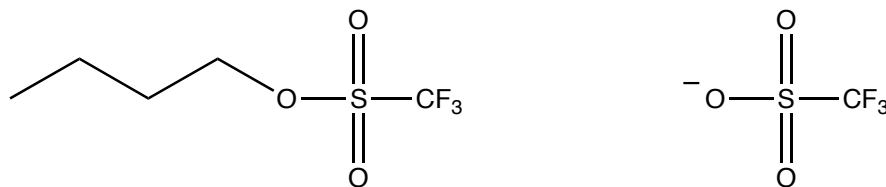


Quiz 3

1. Determine the products of the following reactions.



2. Why are sulfonate esters, like triflate, good leaving groups?



The negative charge is distributed via resonance over three oxygen atoms. The electronegative fluorine atoms also help to draw electron density off the oxygen atoms. Since the negative charge is distributed over so many atoms, triflate (sulfonates) are weak bases, which means that they are good leaving groups.