(37) **Today**

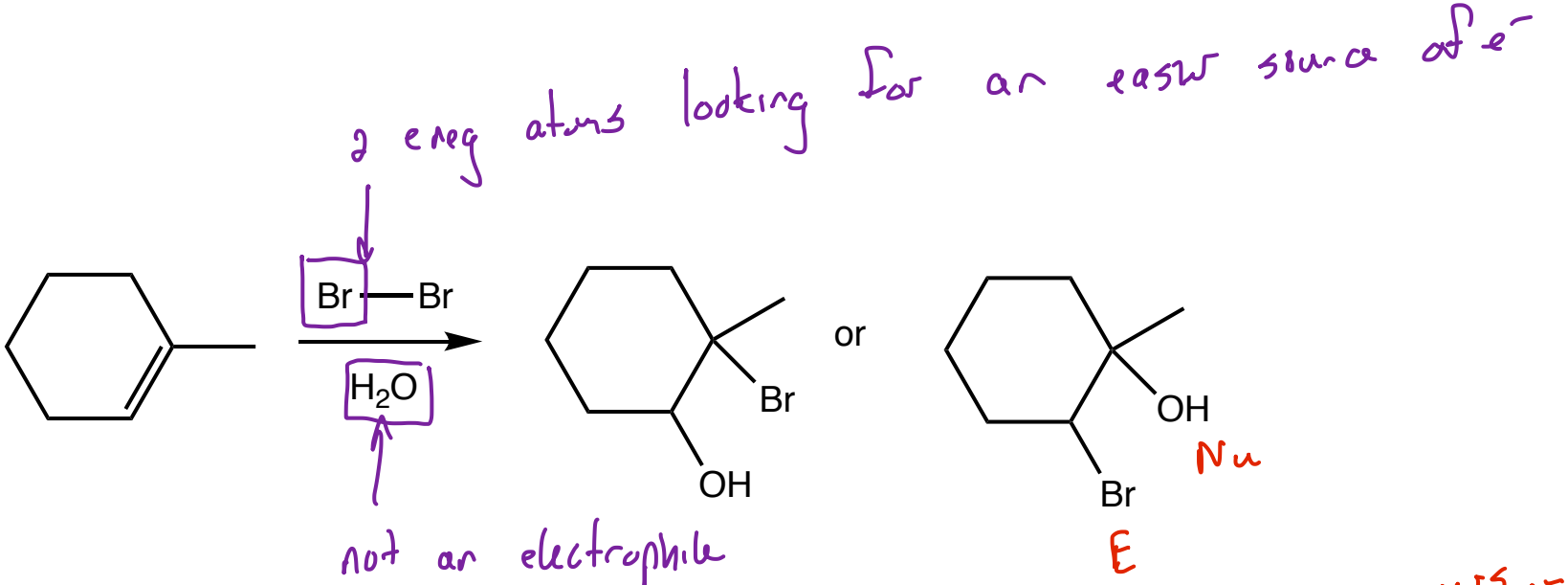
Section 8.2 and 8.3 Halogenation and
Halohyrins: Bromohydrin formation

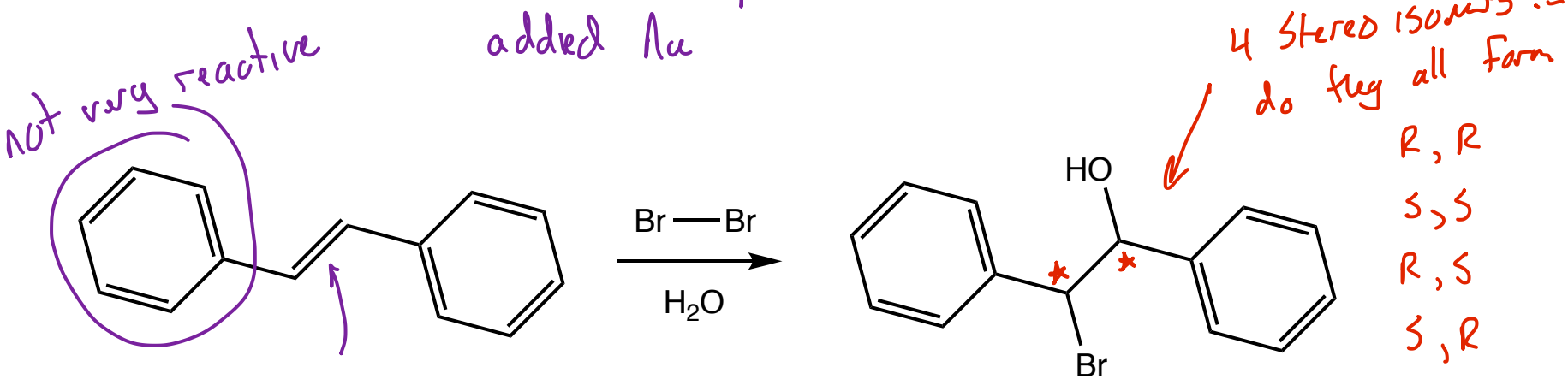
Review

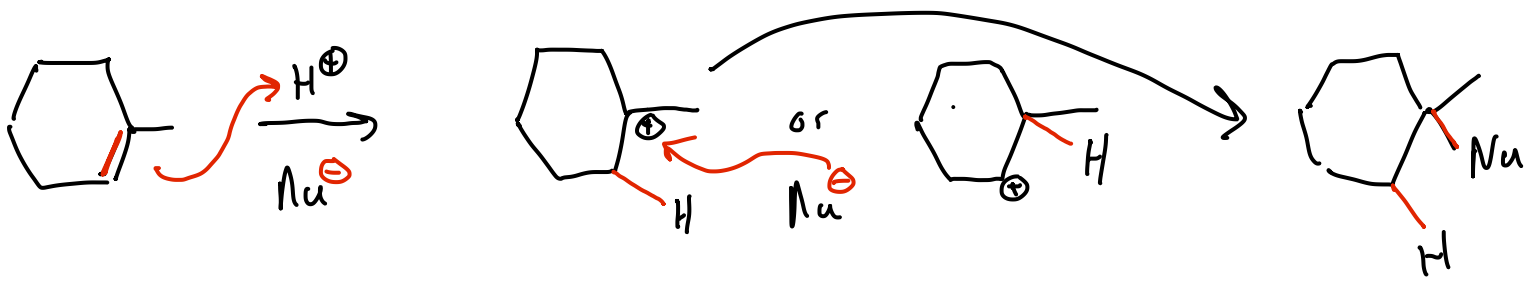
Final Is Scheduled for Monday, December 19 from 12:20 to 2:20

The final will be approximately 1 page each from tests 1, 2, and 3, and 1.5 to 2 pages on alkene nomenclature, stability, carbocations, and electrophilic addition.

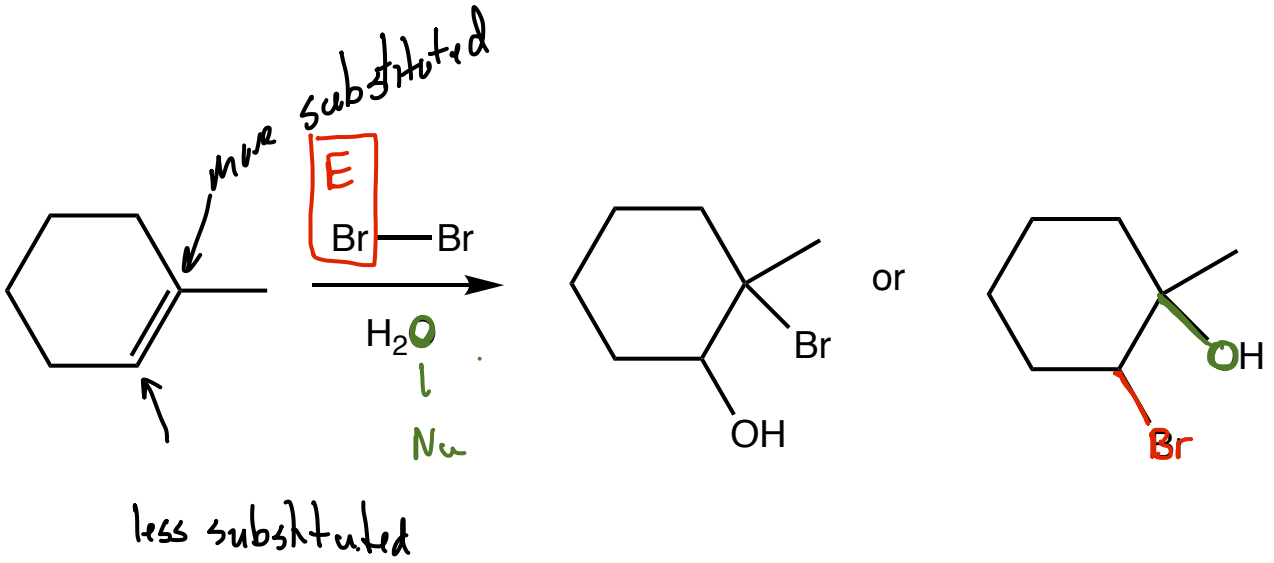
The Reactions





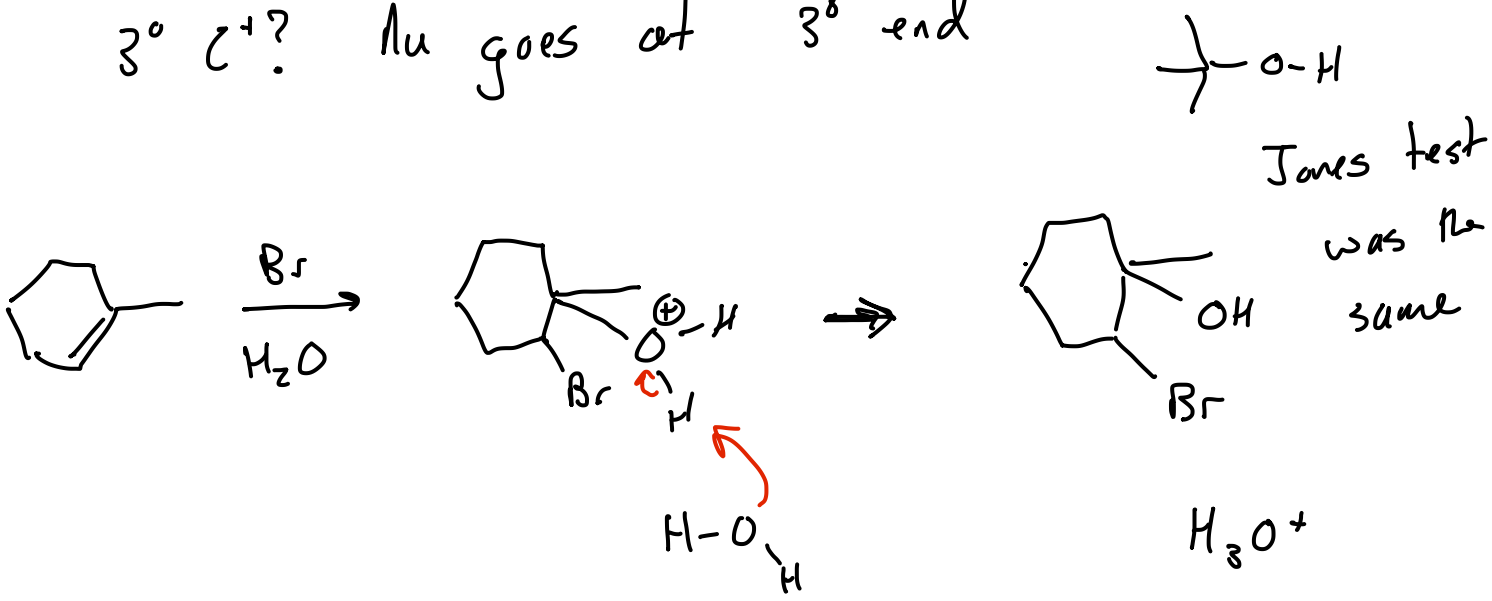


3° C⁺ is easier to make



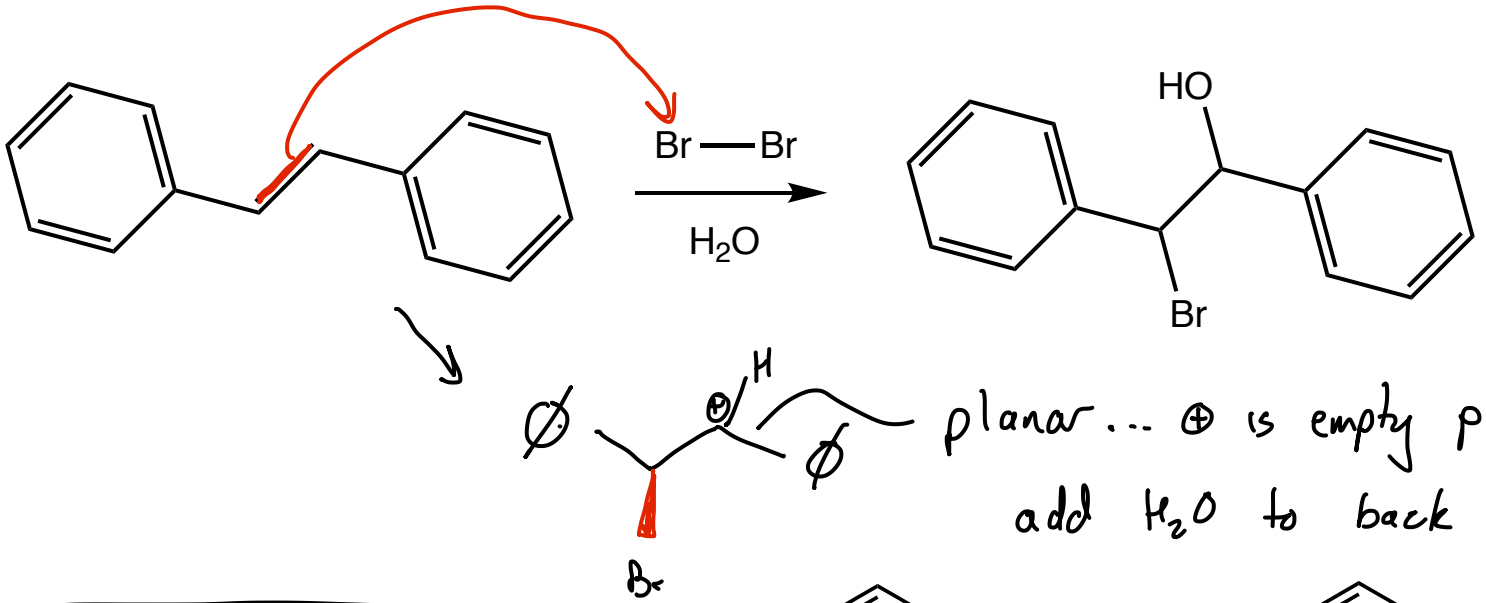
Markovnikov rule ... Nu goes the end of the db that is better at forming a $C^{\oplus}$

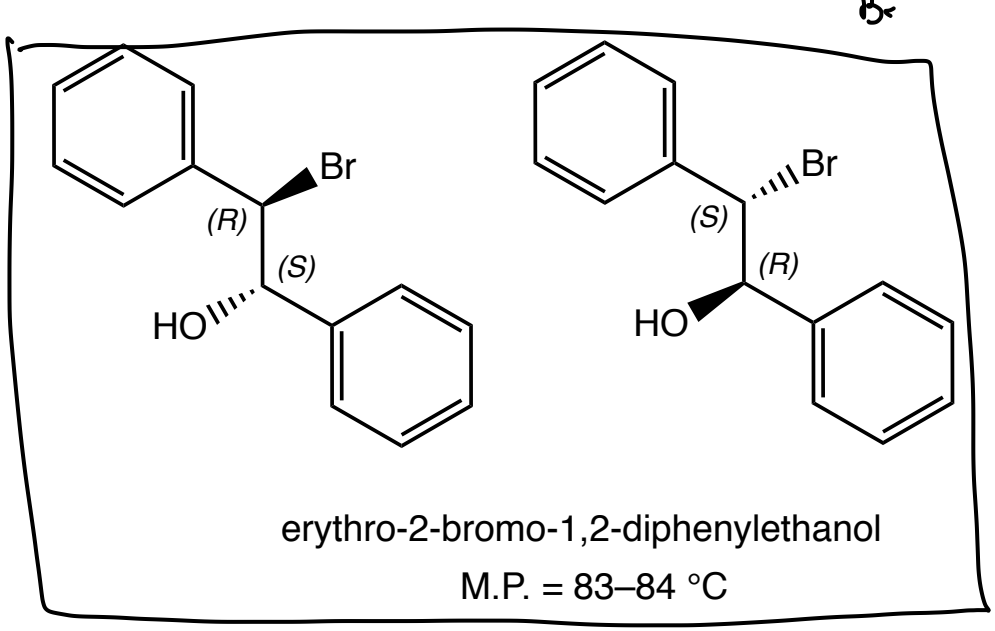
2° C^{+}? 3° C^{+}? Nu goes at 3° end



Jones test was the same

Stereochemistry

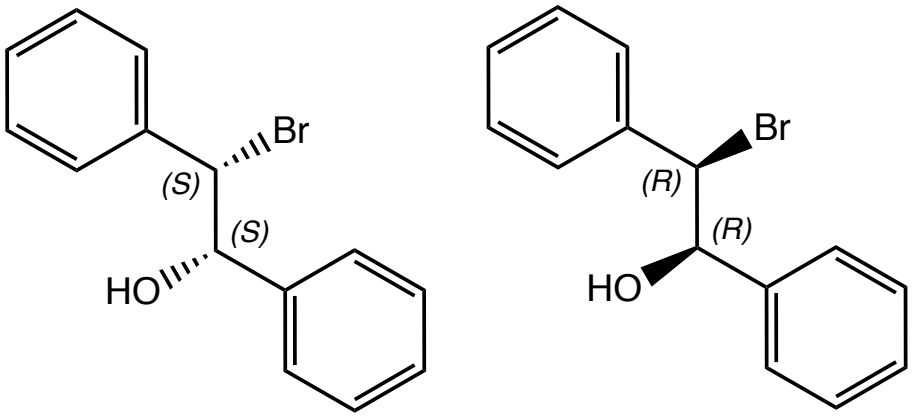




erythro-2-bromo-1,2-diphenylethanol

M.P. = 83–84 °C

products

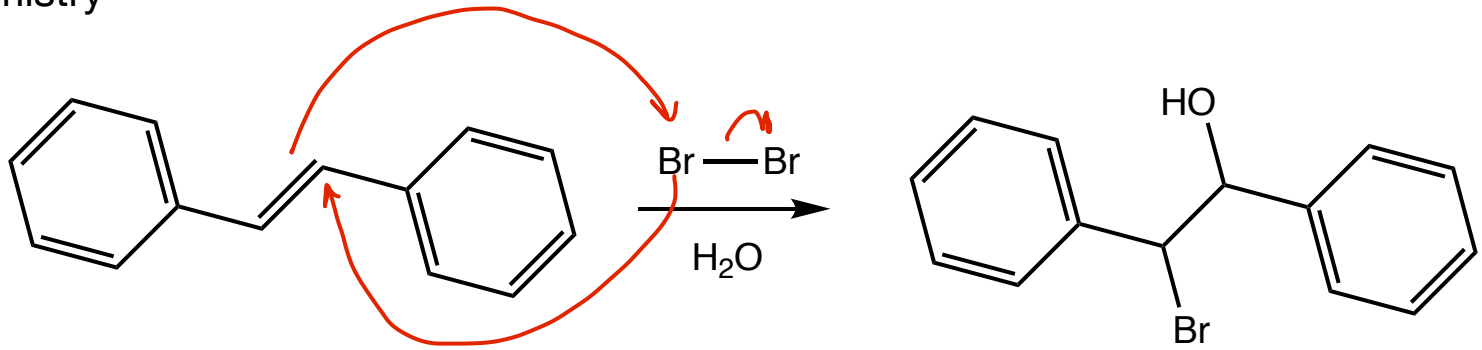


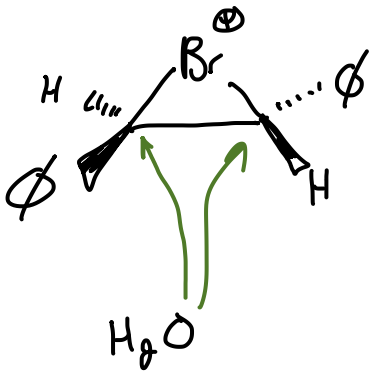
threo-2-bromo-1,2-diphenylethanol

M.P. = 51–52 °C

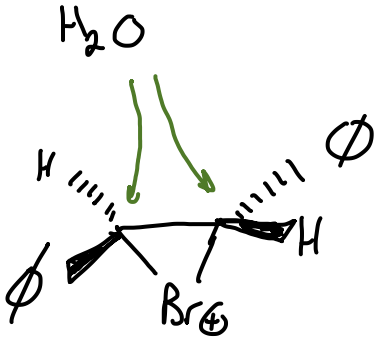
not produced

Stereochemistry



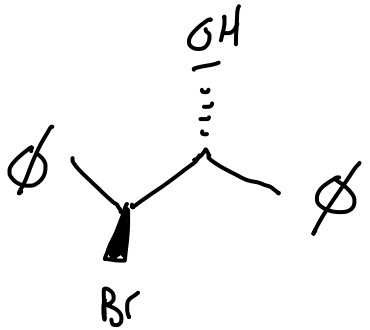


or

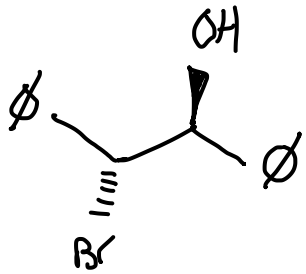


$Br^{\ominus}$

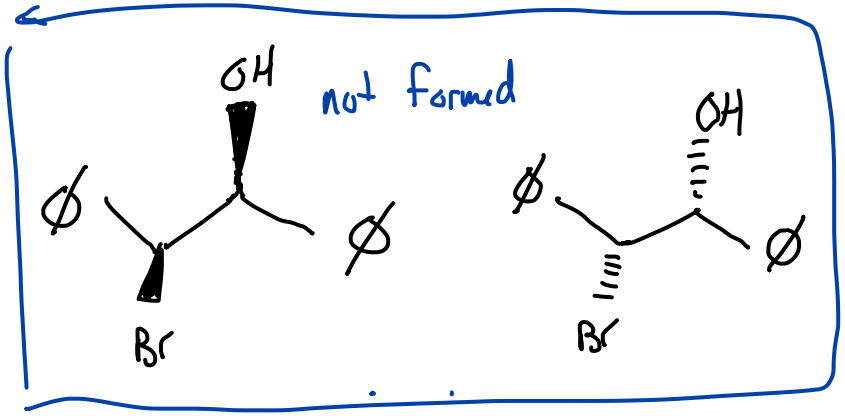
$-H^{+}$



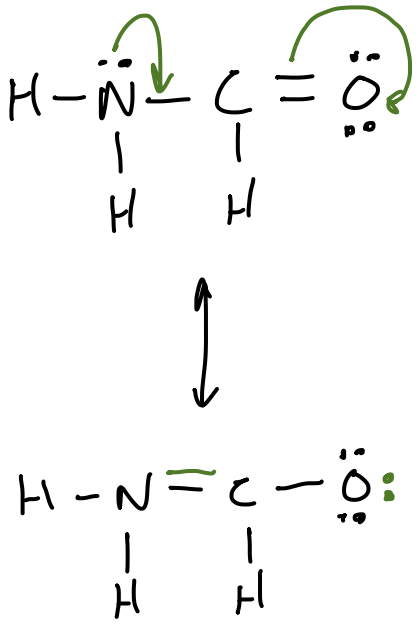
or



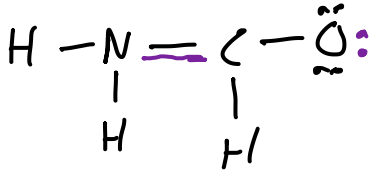
H_2O can't come in from down here... Br is blocking the way



ϕ = 



switch the end of the db & the lp e⁻'s
beginning of db stays in the middle where it originally was

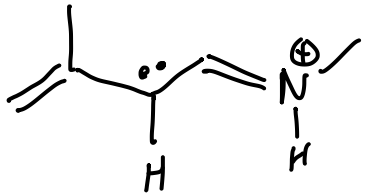


$FC_C = 4 - (2 + 3)$

$= -1$

$FC_N = 5 - (0 + 4)$

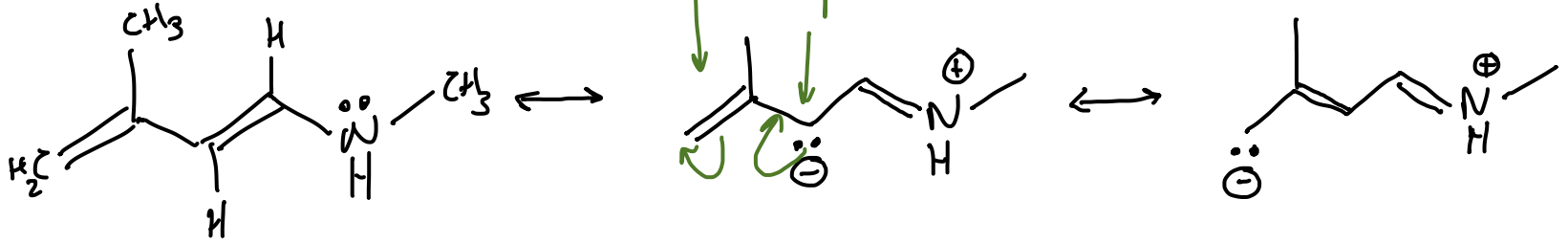
$= 1$



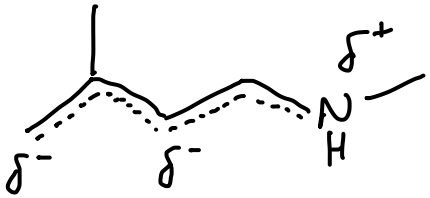
π bond

lp e⁻

resonance contributors



molecule most strongly resembles the most normal looking Lewis structure



resonance hybrid best model of actual molecule